

Stoichiometry Stumper 1 Answer

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Stoichiometry Stumper 1 Answer

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stoichiometry stumper #1 by Amy Hinds - Prezi

Stoichiometry Stumper #1 You are a forensic investigator. You are investigating a murder involving poison. The victim was poisoned with a compound called di-chloro benzene whose formula is C6H4Cl2.

Solved: Stoichiometry Stumper #1 You Are A Forensic Scientist ...

Stoichiometry Stumper #1 You are a forensic scientist. You are investigating a murder involving poison. The victim was poisoned with a compound called di-chloro benzene whose formula is C6H4Cl2. Autopsy results show that the victim's body contained about 31 g of the poison, but the actual amount could have been slightly

Stoichiometry Stumper #1 - Tumwater School District

Stoichiometry Stumper 1 Answer Forensic Author: electionsdev.calmatters.org-2020-10-22T00:00:00+00:01 Subject: Stoichiometry Stumper 1 Answer Forensic Keywords: stoichiometry, stumper, 1, answer, forensic Created Date: 10/22/2020 5:50:48 PM

Stoichiometry Stumper 1 Answer Forensic

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If 1 mole of amus is the same mass as 1 gram, and 1 hydrogen atom has a mass of 1 amu, then 1 mole of hydrogen atoms would have a mass of 1 &#e; ... stoichiometry stumper 1 answer forensic - Bing Balance the equation Math Math Explanation Astronauts died as they could only get rid of 2,750.625 grams of carbon dioxide and needed to get rid of 3,000 grams of carbon dioxide.

Stoichiometry Stumper 1 Answer Forensic

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3 1 1.47 30.0 g 122.55KCl 1 mol KClO mol KCl 3 g KCl 1 3KCl mol KClOmol = g KClO49.3 74.55 2 2.73 1KClO 3 mol KCl molg KClO KCl 3 2 mol KCl O 3 = g KCl 122.55 1 1.66 74.55 g2 KCl KCl 42.7 g Fe mole Fe mol2 Fe 2O 3 g Fe Femol FeO 3 g Fe 2O 3 = g Fe 2O 3 55.85 4 1 61.0 17.0 g O 2 1 mol 2O 2 mol Fe 2O 3 g O32.00 3 mol O 2 g Fe 2O 3 g Fe 2O 3 = 1 ...

stoichiometry 1 worksheet and key - Saddleback College

This was a problem in my chemistry review packet. It's the only one I don't know I don't know how to do. I've been working on it for hours. A mixture of methane (CH4) and ethane (C2H6) is burned. The two combustion reactions require a total of 179.959 grams of oxygen, and they produce a total of 226.473 grams of products. From this information, calculate the masses of methane and ethane in the ...

Stoichiometry Stumper? | Yahoo Answers

1 mol O x mol H O 22.4 L 2O 2 2 2 2 2 D. Find the mass of tristearin required to produce 55.56 moles of water (about 1 liter of liquid water). 57 110 6 57 110 6 57 110 6 2 57 110 6 57 110 6 2 899g C H O 1 mol C H O 890 g C H O 110 mol H O 2 mol C H O x 55.6 mol Answers: 254A. 9.9 x 10 atoms Mn 4C. 33.2 mol Mn 3 O 4 5A. 1168 L O 2 5C. 0.675 ...

Stoichiometry: Problem Sheet 1

The Answer Our company does not have enough funds to produce the calcium nitrate needed to make the required amount of the final drug. Our company will have to sell the rights to the drug to a more well-funded company because we only have \$40,000 and the amount needed is about

Stoichiometry Stumper #3 by Jadine Lee - Prezi

Stoichiometry Worksheet #1 Answers 1. Given the following equation: 2 C4H10 + 13 O2--> 8 CO2 + 10 H2O, show what the following molar ratios should be. a. C4H10 / O2 b. O2 / CO2 c. O2 / H2O d. C4H10 / CO2 e. C4H10 / H2O 2. Given the following equation: 2 KClO3--> 2 KCl + 3 O2 a. How many moles of O2 can be produced by ...

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Stoichiometry Level 1, Level 2, Level 3 Answers. Basic Skills in Chemistry Workbook Stoichiometry Level 1, Level 2, Level 3 Answers: Level 1: 1. 55.7 g 2. .982 g 3. 164 g 4. 64.2 g 5. 19.4 g. Found: 12 Mar 2020 | Rating: 84/100. Practice Test Ch 3 Stoichiometry Name Per. Remember it is a MC test, use the answers ...

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