

Chapter 10 Chemical Quantities Answer Key

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Chapter 10 Chemical Quantities. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. claire_kutchka. mrs. wright. Terms in this set (38) standard temperature and pressure. 0 C and 101.3 kPa. the volume occupied by a mole of any gas at STP (22.4 L) molar volume.

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To determine a compound/element's molar mass: 1) Make a list of all the elements and the number of that specific element. 2) Use the periodic table to write down the mass for each element. 3) After you have determined the mass and number of each element, you will multiply the mass by the number of atoms. 4) Once you have multiplied, you must add all the numbers together to get your final answer.

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Chapter 10 - Chemical Quantities Section 10.1 - The Mole: A Measurement of Matter You often measure the amount of something by count, by mass, or by volume. A mole (mol) of a substance is 6.02×10^{23} representative particles of that substance.

Chemical Quantities Section 10 1 The Mole A Measurement Of ...

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Chapter 10: Chemical Quantities

CHAPTER 10: Chemical Quantities BASICS: • The basic unit that is used to determine the amount of a chemical substance is called a mole • A mole(mol) of a substance is equivalent to 6.02×10^{23} particles of that substance • The mole was founded by a scientist named Avagadro, and he decided to use the

CHAPTER 10: Chemical Quantities

Chapter 10 "Chemical Quantities" Vocab. the SI unit representing 6.02×10^{23} representative particles of a substance. the temperature and pressure at which one mole of gas occupies a volume of 22.4 L. equal volumes of gases at the same temperature and pressure contain equal numbers of particles.

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Chapter 10 Chemical Quantities91 SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287-296) This section defines the mole and explains how the mole is used to measure matter. It also teaches you how to calculate the mass of a mole of any substance. Measuring Matter (pages 287-289) 1.

SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287-296)

Section 10.1 - The Mole: A Measurement of Matter. You often measure the amount of something by count, by mass, or by volume. A mole (mol) of a substance is 6.02×10^{23} representative particles of that substance. 6.02×10^{23} is called Avogadro's number. 1 mole = 6.02×10^{23} representative particles

Chapter 10 - Chemical Quantities

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Chapter 10 Chemical Quantities Answer Key Page 247

Chemical Quantities THE MOLE AND QUANTIFYING MATTER 10.1 The Mole: A ... After reading Lesson 10.1, answer the following questions. ... The quantities 1 mol and 22.4 L can be used in conversion factors that change moles to volume and volume to moles at STP.

Chemical Quantities - AP Biology

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